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Basic Chemistry Part 1 (1-25) 1) When a beam of white rays is dispersed by a prism which
colour will be refracted to a larger extent? Answer: Violet. 2) The modern atomic mass unit is
based on ____ Answer: Carbon-12. 3) What do you mean by photoelectric effect?

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Question & Answers (0) The relationship between the coefficients of the components of a balanced chemical is referred to as the: a) tracheotomy b) symmetry c) stoichiometry d) numerology e)...

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This hub examines the answers to some of the top chemistry-related science questions I have been asked by my students in our science lessons. 1. What Is an Acid? Put simply, an acid is any substance that has pH of less than 7. The pH scale is used to measure how acid or alkali a substance is: 0-3 = strong acid (UI turns red)

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1. What is the first element on the periodic table? 2. What is the centre of an atom called? 3. True or false? Acids have a pH level below 7. 4. What is the main gas found in the air we breathe?

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Boy (8-10) in Chemistry Lab. Charles Thatcher, Getty Images If there had been a million dollars on the line for knowing all the answers, you would have just missed the prize. It's OK to miss a question or two. That just shows you have other interests! You have a good command of chemistry.

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Answers: Nickel, cobalt and gadolinium. Helium; Xenon; Sir Humphry Davy; Van der Waals forces (or bonds) (a)Antimony, (b)Potassium, (c)Silver, (d)Tungsten, and (e)Californium? Laughing gas; Dmitri Mendeleev; Antifreeze; Linus Pauling (the only person to be awarded two unshared Nobel Prizes) Vitamin C; Copper and zinc; Copper and tin; Lenz's Law; J; Tungsten; Chemistry Quiz II

Chemistry Quiz, Chemistry Quiz Questions and Answers

If you study physics, you should be able to explain why the sky is blue. If biology is your thing, you ought to be able to answer where babies come from. Chemistry doesn't have any great standard questions, but there are some everyday phenomena you should be able to explain. 01

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O Level Chemistry 5090 Past Papers March, May & November ...
to calculate reacting masses and product masses. Example 1 – reactant mass Question. When 12 g of carbon is burned in air, 44 g of carbon dioxide is produced.

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